15% solution by mass means

Mass of solution = 100g

Mass of solute = 15g

Number of moles of  $H_2SO_4 = 15/98$ 

$$H_2SO_4$$
 Density =  $\frac{Mass}{Volume}$ 

$$1 \times 2 + 32 + 16 \times 4 = 98$$
  $1.2 = \frac{100}{\text{Volume}}$   $1.2 = \frac{100}{\text{Volume}} \times 10 = \frac{1000}{12} - 0.012 \text{ml}$ 

$$\frac{100}{1.2}$$
 solution contain =  $\frac{15}{98}$ 

1 solution contain = 
$$\frac{15}{98} \times \frac{10}{100}$$

1000 solution contain = 
$$\frac{15}{98} \times \frac{1.2}{100} \times 1000$$

$$= 1.8 M.$$

#### 2.4 MOLALITY

Number of moles of the solute present in 1000 g solvent. It is represented by 'M'.

#### 2.5 NORMALITY

Number of gram equivalents of a solute in 1000 ml solution. It is represented by 'N'.

Number of gram equivalents = 
$$\frac{\text{Mass in gram}}{\text{equivalent Mass}}$$

#### FORMULAE FOR EQUIVALENT MASS

Equivalent mass of acid = 
$$\frac{\text{Molecular Mass of Acid}}{\text{Basicity}}$$

Equivalent mass of base = 
$$\frac{\text{Molecular Mass of Base}}{\text{Acidity}}$$

Equivalent mass of salt = 
$$\frac{\text{Formula mass of salt}}{\text{twevalency of cation}}$$

#### 2.6 MOLE FRACTION

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# **Class 12th Notes Chemistry Chapter Wise Notes**

**Annelies Wilder-Smith** 

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