

Station #1: Classifying Reactions

- | | | |
|-----|---|--------------------|
| 1. | $\text{PbCl}_2 + \text{AgNO}_3 \rightarrow \text{Pb}(\text{NO}_3)_2 + \text{AgCl}$ | Double Replacement |
| 2. | $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$ | Synthesis |
| 3. | $\text{AlCl}_3 + \text{Na}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{NaCl}$ | Double Replacement |
| 4. | $\text{Zn} + \text{S} \rightarrow \text{ZnS}$ | Synthesis |
| 5. | $\text{Al}_2(\text{SO}_4)_3 + \text{BaCl}_2 \rightarrow \text{BaSO}_4 + \text{AlCl}_3$ | Double Replacement |
| 6. | $\text{Al}_2\text{S}_3 \rightarrow \text{Al} + \text{S}$ | Decomposition |
| 7. | $\text{H}_2\text{SO}_4 + \text{Fe} \rightarrow \text{H}_2 + \text{FeSO}_4$ | Single Replacement |
| 8. | $\text{C}_{12}\text{H}_{22}\text{O}_{11} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$ | Combustion |
| 9. | $\text{Mg}(\text{OH})_2 + \text{H}_2\text{SO}_4 \rightarrow \text{MgSO}_4 + \text{H}_2\text{O}$ | Double Replacement |
| 10. | $\text{NaOH} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu}(\text{OH})_2$ | Double Replacement |
| 11. | $\text{C}_6\text{H}_{12} + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{CO}_2$ | Combustion |
| 12. | $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ | Combustion |
| 13. | $\text{Mg}_3(\text{PO}_4)_2 + \text{H}_2 \rightarrow \text{Mg} + \text{H}_3\text{PO}_4$ | Single Replacement |
| 14. | $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$ | Decomposition |
| 15. | $\text{Cl}_2 + \text{KBr} \rightarrow \text{KCl} + \text{Br}_2$ | Single Replacement |

Station #2: Balancing Reactions

1. $1 \text{ Na}_3\text{PO}_4 + 3 \text{ KOH} \rightarrow 3 \text{ NaOH} + 1 \text{ K}_3\text{PO}_4$
2. $2 \text{ RbNO}_3 + 1 \text{ BeF}_2 \rightarrow 1 \text{ Be}(\text{NO}_3)_2 + 2 \text{ RbF}$
3. $1 \text{ Pb}(\text{OH})_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$
4. $2 \text{ Na}_3\text{PO}_4 + 3 \text{ CaCl}_2 \rightarrow 6 \text{ NaCl} + 1 \text{ Ca}_3(\text{PO}_4)_2$
5. $2 \text{ NH}_3 + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 (\text{NH}_4)_2\text{SO}_4$

Station #3: Single Displacement Equations

For each of the following SD reactions, predict the products that are formed and balance the equation. If no reaction occurs, write "NR"



Classifying Chemical Reactions Answers Physical Science If8767

Eve Hartman, Wendy Meshbesher



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Exploring Chemical Reactions Nigel Saunders,2007 *The Basics of Chemical Reactions* Krista West,2013-12-15
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Understanding Chemistry Robert J. Ouellette,1987 **Chemicals & Reactions** Jon Richards,2004-08 Chemical reactions are happening around us and inside us every day Every time we boil water wash our hands or stir sugar into cocoa a chemical reaction is taking place Book jacket **Reactions** Theodore W. Gray,2017 The book begins with a brief recap of

elements and molecules and then goes on to explain important concepts that characterize a chemical reaction including Energy Entropy and Time. It is then organized by type of reaction including Combustion Photosynthesis Respiration Oxidation and Fermentation. A special section dedicated to chemical cycles includes The Carbon Cycle The Iron Cycle and The Lime Cycle. *Chemical Reactions* Eve Hartman, Wendy Meshbesher, 2009-07-10 *Chemical Reactions* Nigel Saunders, 2007. This book explains what chemical reactions are what elements atoms and molecules are and what happens to them during chemical reactions. **Chemical Reactions** Louise Spilsbury, Richard Spilsbury, 2007. Readers will learn what chemical reactions are how they work what changes happen during reactions and how we can stop reactions. *Encyclopedia of Chemical Reactions* Carl Alfred Jacobson, 1946.

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