

where p is pressure and v is specific volume, or the volume displaced per unit mass evaluated at the inlet or exit.

A **property** of a system is any observable characteristic of the system. The **state** of a system is defined by specifying the minimum set of independent properties. The most common thermodynamic properties are temperature T , pressure p , and specific volume v or density ρ . Additional thermodynamic properties include entropy, stored forms of energy, and enthalpy.

Frequently, thermodynamic properties combine to form other properties. **Enthalpy** h is an important property that includes internal energy and flow work and is defined as

$$h = u + pv \quad (4)$$

where u is the internal energy per unit mass.

Each property in a given state has only one definite value, and any property always has the same value for a given state, regardless of how the substance arrived at that state.

A **process** is a change in state that can be defined as any change in the properties of a system. A process is described by specifying the initial and final equilibrium states, the path (if identifiable), and the interactions that take place across system boundaries during the process.

A **cycle** is a process or a series of processes wherein the initial and final states of the system are identical. Therefore, at the conclusion of a cycle, all the properties have the same value they had at the beginning. Refrigerant circulating in a closed system undergoes a cycle.

A **pure substance** has a homogeneous and invariable chemical composition. It can exist in more than one phase, but the chemical composition is the same in all phases.

If a substance is liquid at the saturation temperature and pressure, it is called a **saturated liquid**. If the temperature of the liquid is lower than the saturation temperature for the existing pressure, it is called either a **subcooled liquid** (the temperature is lower than the saturation temperature for the given pressure) or a **compressed liquid** (the pressure is greater than the saturation pressure for the given temperature).

When a substance exists as part liquid and part vapor at the saturation temperature, its **quality** is defined as the ratio of the mass of vapor to the total mass. Quality has meaning only when the substance is saturated (i.e., at saturation pressure and temperature). Pressure and temperature of saturated substances are not independent properties.

If a substance exists as a vapor at saturation temperature and pressure, it is called a **saturated vapor**. (Sometimes the term **dry saturated vapor** is used to emphasize that the quality is 100%.) When the vapor is at a temperature greater than the saturation temperature, it is a **superheated vapor**. Pressure and temperature of a superheated vapor are independent properties, because the temperature can increase while pressure remains constant. Gases such as air at room temperature and pressure are highly superheated vapors.

1.3 FIRST LAW OF THERMODYNAMICS

The first law of thermodynamics is often called the **law of conservation of energy**. The following form of the first-law equation is valid only in the absence of a nuclear or chemical reaction.

Based on the first law or the law of conservation of energy, for any system, open or closed, there is an energy balance as

$$\left[\begin{array}{c} \text{Net amount of energy} \\ \text{added to system} \end{array} \right] = \left[\begin{array}{c} \text{Net increase of stored} \\ \text{energy in system} \end{array} \right]$$

or

$$[\text{Energy in}] - [\text{Energy out}] = [\text{Increase of stored energy in system}]$$

Figure 1 illustrates energy flows into and out of a thermodynamic system. For the general case of multiple mass flows with uniform properties in and out of the system, the energy balance can be written

$$\begin{aligned} & \sum m_{in} \left(u + pv + \frac{V^2}{2} + gz \right)_{in} \\ & - \sum m_{out} \left(u + pv + \frac{V^2}{2} + gz \right)_{out} + Q - W \\ & = \left[m_f \left(u + \frac{V^2}{2} + gz \right)_f - m_i \left(u + \frac{V^2}{2} + gz \right)_i \right]_{system} \end{aligned} \quad (5)$$

where subscripts i and f refer to the initial and final states, respectively.

Nearly all important engineering processes are commonly modeled as steady-flow processes. Steady flow signifies that all quantities associated with the system do not vary with time. Consequently,

$$\begin{aligned} & \sum_{\text{all streams entering}} m \left(h + \frac{V^2}{2} + gz \right) \\ & - \sum_{\text{all streams leaving}} m \left(h + \frac{V^2}{2} + gz \right) + Q - W = 0 \end{aligned} \quad (6)$$

where $h = u + pv$ as described in Equation (4).

A second common application is the closed stationary system for which the first law equation reduces to

$$Q - W = [m(u_f - u_i)]_{system} \quad (7)$$

1.4 SECOND LAW OF THERMODYNAMICS

The second law of thermodynamics differentiates and quantifies processes that only proceed in a certain direction (irreversible) from those that are reversible. The second law may be described in several ways. One method uses the concept of entropy flow in an open system and the irreversibility associated with the process. The concept of irreversibility provides added insight into the operation of cycles. For example, the larger the irreversibility in a refrigeration cycle operating with a given refrigeration load between two fixed temperature levels, the larger the amount of work required to operate the cycle. Irreversibilities include pressure drops in lines and heat exchangers, heat transfer between fluids of different temperature, and mechanical friction. Reducing total irreversibility in a cycle improves cycle performance. In the limit of no irreversibilities, a cycle attains its maximum ideal efficiency.

In an open system, the second law of thermodynamics can be described in terms of entropy as

$$dS_{system} = \frac{\delta Q}{T} + \delta m_e s_e - \delta m_s s_s + dI \quad (8)$$

where

dS_{system} = total change within system in time dt during process

$\delta m_e s_e$ = entropy increase caused by mass entering (incoming)

$\delta m_s s_s$ = entropy decrease caused by mass leaving (exiting)

$\delta Q/T$ = entropy change caused by reversible heat transfer between

system and surroundings at temperature T

dI = entropy caused by irreversibilities (always positive)

Equation (8) accounts for all entropy changes in the system. Rearranged, this equation becomes

$$\delta Q = T[(\delta m_e s_e - \delta m_s s_s) + dS_{system} - dI] \quad (9)$$

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