Atomic Target Practice

Solving the Structure of the Atom



Introduction

Rutherford scattering is one of the most famous experiments of all time. More than 25 years after conducting the experiment, Lanest Rutherford described the results this way:

"It was about as credible as if you had fired a 15-inch shell at a piece of timus paper and it came back and hit you."

The experiment itself was actually the culmination of a series of experiments, carried out over a five-year period, dealing with the scattering of high-energy alpha particles by various substances. What is "Batherford scattering" and why was it so important?

Concepts

According strengthurse.

· Nucleus

Accumie muso.

Nuclear charge

Background

Farnest Rutherford received the Nobel Prize in Chemistry in 1908 for his investigations into the disintegration of the elements as a result of radioactive decay. Among the products of the radioactive decay of elements are alpha particles—small, positively charged, high-energy particles. In trying to learn more about the nature of alpha particles, Rutherford and his co-workers, Hans Geiger and Ernest Marsden, began studying what happened when a narrow brancot alpha particles was directed at a thin piece of metal foil. Alpha particles are a type of nuclear radiation, traveling at about 1/10 the speed of light. As espected for such high-energy particles, most of the particles penetrated the thin metal foil and were detected on the other side. What was one spected was that a few—a very few, to be some—of the alpha particles were actually reflected back toward the source, having been "seattered" or bent due to their encounters with the metal arous in the foil target. The number of alpha particles that were reflected back depended on the arouse mass of the metal. Gold amous, having the highest arouse of the metals studied, gave the largest amount of so-called "back scattering."

Ratherford's scattering experiments have been described as a "black bus" experiment. The prospecties of the alpha particles, their mass, charge, speed, etc., were at least partially understand. The atoms making up the target, however, presented Batherford with a black bos—the structure of the atom was not known at the time. In order to explain the results of the scattering experiments. Ratherford had to unlock the black bos, that is, he had to solve the structure of the atom. In 1911 Rutherford proposed the following model for the structure of the atom.

- Most of the mass of the atom is concentrated in a very small, dense central area, later called the nucleus, which is about 1/100,000 the diameter of the atom.
- The rost of the store is apparently "compty space." (Most of the alpha particles traveled straight through the metal foil, as
 if mothing were in their path.)
- The central, dense care of the atom is positively charged, with the nuclear charge equal to about one-half the atomic
 trass. (is alpha particles randomly struck the gold target, a few approached the modern of an atom local on. The positively
 charged alpha particles were strongly repelled by the modern charge and "revolud" or bounced back to the source.)

It is not practical to recreate Rutherford's original scattering experiments in the high school laboratory. Some idea of the challenge that faced Rutherford and his co-workers can be gained from the following "black box" activity using marbles and an unseen, unknown target.

Atomic Target Practice Lab Answers

Gardner Fox, Al Feldstein

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