

Atomic Structure

1. ATOM & MOLECULES

- (a) The smallest particle of a matter that takes part in a chemical reaction is called an atom. The atom of all gases except those of noble gases, cannot exist in free state. These exist in molecular form. The molecules of hydrogen, nitrogen, oxygen and halogens are diatomic (H_2 , N_2). Phosphorus molecule is tetratomic and that of sulphur is octa atomic.
- (b) The smallest particle of a matter that can exist in free state in nature, is known as a molecule.
- (c) Some molecules are composed of homoatomic atom, e.g., H_2 , O_2 , N_2 , Cl_2 , O_3 etc., while the molecules of compounds are made up of two or more heteroatomic atoms e.g., HCl , $NaOH$, HNO_3 , $CaCO_3$, etc.

2. DALTON'S ATOMIC THEORY

The concepts put forward by John Dalton regarding the composition of matter are known as Dalton's atomic theory. Its important points are as follows.

- (a) Every matter is composed of very minute particles, called atoms that take part in chemical reactions.
- (b) Atoms cannot be further subdivided.
- (c) The atoms of different elements differ from each other in their properties and masses, while the atoms of the same element are identical in all respects.
- (d) The atoms of different elements can combine in simple ratio to form compounds. The masses of combining elements represent the masses of combining atoms.
- (e) Atom can neither be created nor destroyed.

2.1 Modern Concept :

Many of the concepts of Dalton's atomic theory cannot be explained. Therefore, foundation of modern atomic theory was laid down by the end of nineteenth century. The modern theory is substantiated by the existence of isotopes, radioactive disintegration, etc. The important points of the modern atomic theory are as follows.

- (a) Prof. Henri Becquerel discovered the phenomenon of radioactivity and found that an atom is divisible.
- (b) An atom is mainly composed of three fundamental particles, viz. electron, proton and neutron.
- (c) Apart from the aforesaid three fundamental particles, many others have also been identified, viz. positron, meson, neutrino, antiproton, etc.
- (d) Soddy discovered the existence of isotopes, which were atoms of the same element having different masses. For example, protium, deuterium and tritium are atoms of hydrogen having atomic masses 1, 2 and 3 a.m.u. respectively.
- (e) Atoms having same mass may have different atomic numbers. These are known as isobars. For example, $^{40}_{18}Ar$ and $^{40}_{20}Ca$.
- (f) Atoms of elements combines to form molecules.
- (g) It is not necessary that the atoms should combine in simple ratio for the formation of compounds. The atoms in non-stoichiometric compounds are not present in simple ratio. For example, in ferrous sulphide crystals, iron and sulphur atoms are present in the ratio of 0.86 : 1.00.
- (h) Atoms participate in chemical reactions.

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