

Stoichiometry Worksheet

I. ONE REACTANT IN EXCESS – COMBUSTION

You burn an excess of butane (C₄H₁₀) in a limited amount of oxygen (O₂). If 67.2 L of CO₂ gas is formed, what volume of oxygen gas was present at the start of the reaction?

STEP 1: Balance the reaction:



STEP 2: Use a molar conversion to convert quantity of product given to moles:

$$67.2 \text{ L CO}_2 * \frac{(1 \text{ mol CO}_2)}{(22.4 \text{ L CO}_2)} = 3 \text{ mol CO}_2$$

STEP 3: Use a molar ratio to calculate the moles of reactant that I'm asking about:

$$3 \text{ mol CO}_2 * \frac{(13 \text{ mol O}_2)}{(8 \text{ mol CO}_2)} = 4.875 \text{ mol O}_2$$

STEP 4: Use a molar conversion to convert moles of reactant to the units I ask for:

$$4.875 \text{ mol O}_2 \frac{(22.4 \text{ L O}_2)}{(1 \text{ mol O}_2)} = 109.2 \text{ L O}_2 = \boxed{109 \text{ L O}_2}$$

II. ONE REACTANT IN EXCESS – SINGLE DISPLACEMENT

14 g of lithium is combined with excess K₂SO₄. What is the mass of the new ionic compound that is formed? How do you know that this reaction will take place?

STEP 1: Balance the reaction:



Reaction occurs because Li is higher up than K on the activity series.

STEP 2: Use a molar conversion to convert quantity of reactant given to moles:

$$14 \text{ g Li} * \frac{(1 \text{ mol Li})}{(7 \text{ g Li})} = 2 \text{ mol Li}$$

STEP 3: Use a molar ratio to calculate moles of product that I'm asking about:

$$2 \text{ mol Li} * \frac{(1 \text{ mol Li}_2\text{SO}_4)}{(2 \text{ mol Li})} = 1 \text{ mol Li}_2\text{SO}_4$$

STEP 4: Use a molar conversion to convert moles of product to units that I ask for:

$$1 \text{ mol Li}_2\text{SO}_4 * \frac{(110 \text{ g Li}_2\text{SO}_4)}{(1 \text{ mol Li}_2\text{SO}_4)} = \boxed{110 \text{ Li}_2\text{SO}_4}$$

Section Quiz Introduction To Stoichiometry Answers

Donada Peters, Edward I. Peters



Section Quiz Introduction To Stoichiometry Answers:

Chemtrek Stephen Thompson,1990 Introduction to Chemical Processes: Principles, Analysis, Synthesis Regina M. Murphy,2007 Introduction to Chemical Processes Principles Analysis Synthesis enhances student understanding of the connection between the chemistry and the process Users will find strong coverage of chemistry gain a solid understanding of what chemical processes do convert raw materials into useful products using energy and other resources and learn about the ways in which chemical engineers make decisions and balance constraints to come up with new processes and products The author presents material and energy balances as tools to achieve a real goal workable economical and safe chemical processes and products Loaded with intriguing pedagogy this text is essential to a students first course in Chemical Engineering Additional resources intended to guide users are also available as package options such as ChemSkill Builder

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