

## What's the Perfect Concentration of Kool-Aid???

**Purpose:** Determine the concentration (molarity) of properly made Kool-Aid.

**Method:** Make three solutions of Kool-Aid with different concentrations and taste them to decide which is the correct concentration. You will prepare 0.1 L of each of the following Kool-Aid solutions: 0.1 M, 0.4 M, and 0.7 M.

**Materials:**

|                 |                                     |
|-----------------|-------------------------------------|
| Kool-Aid Powder | Popsicle sticks (to stir solutions) |
| Water           | Balance                             |
| Plastic cups    | Ruler                               |

**Procedure:**

1. Calculate how much solid Kool-Aid you will need to make 0.1 L of each solution. (Hint 1: Kool-Aid is mostly sugar ( $C_6H_{12}O_6$ ), so you can assume that the "molar mass" of Kool-Aid is the same as the molar mass of sugar.) (Hint 2: "What you know" is the volume – 0.1 L and known molarity)

Show all of your calculations below.

The "molar mass" of Kool-Aid:

Mass of Kool-Aid needed for 0.1 M solution:

Mass of Kool-Aid needed for 0.4 M solution:

Mass of Kool-Aid needed for 0.7 M solution:



# Kool Aid Concentration Lab Answers

**JS Bruner**

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