

Determination of K_a for a Weak Acid

Pre Lab Calculations

*Watch the provided video with example calculations before doing these problems.

Instructions: Watch the video, work these problems and upload your work and answers before your scheduled Zoom lab time. Do your best and use the video. ****MUST** show work to receive credit!

1. The pH of a 0.20 M solution of an unknown monoprotic acid, HA, is 2.12. Determine the value of the K_a for this acid. Start each problem by writing a balanced acid ionization (since calculating K_a) in water reaction and setting up an appropriate ICE table.

Step 1: Acid ionization reaction: $HA(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + A^-(aq)$.

Step 2: Set up an ICE table

	HA	H ₂ O	A ⁻	H ₃ O ⁺
I	0.20M		0	0
C	-x		x	x
E	0.19M		$7.6 \times 10^{-3}M$	$7.6 \times 10^{-3}M$

Step 3: Determine the $[H_3O^+]_{eq}$ using the pH.

$$[H_3O^+]_{eq} = 10^{-pH} = 10^{-2.12} = 7.6 \times 10^{-3}M$$

Step 4: Fill in your known values in to the ICE table. Do $C_1V_1 = C_2V_2$ dilutions if necessary, to find concentrations. (Hint: this isn't necessary in question 1, but is necessary if you're mixing your acid with anything else).

Step 5: Write the equilibrium expression using your balanced equation from Step 1. Use equilibrium values in your ICE table to solve for K_a .

$$K_a = \frac{[H_3O^+][A^-]}{[HA]} = \frac{(7.6 \times 10^{-3}M)^2}{0.19M} = 3.04 \times 10^{-4}$$

2. A solution is prepared by adding 25.0 mL of 0.20 M HA (a different monoprotic acid than in question 1) to 50.0 mL of 0.20 M NaA. The pH of this solution is 3.45. Determine the value of K_a for this acid.

Hint: Follow the steps as outlined in problem 1. Use $C_1V_1 = C_2V_2$ to find the new initial concentration of acid after dilution with the NaA solution. The equation you used in question 1

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Table of Contents Determination Of Ka Lab Report Answers

1. Understanding the eBook Determination Of Ka Lab Report Answers
 - The Rise of Digital Reading Determination Of Ka Lab Report Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Determination Of Ka Lab Report Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Determination Of Ka Lab Report Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Determination Of Ka Lab Report Answers

- Personalized Recommendations
- Determination Of Ka Lab Report Answers User Reviews and Ratings
- Determination Of Ka Lab Report Answers and Bestseller Lists
- 5. Accessing Determination Of Ka Lab Report Answers Free and Paid eBooks
 - Determination Of Ka Lab Report Answers Public Domain eBooks
 - Determination Of Ka Lab Report Answers eBook Subscription Services
 - Determination Of Ka Lab Report Answers Budget-Friendly Options
- 6. Navigating Determination Of Ka Lab Report Answers eBook Formats
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 - Determination Of Ka Lab Report Answers Compatibility with Devices
 - Determination Of Ka Lab Report Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Determination Of Ka Lab Report Answers
 - Highlighting and Note-Taking Determination Of Ka Lab Report Answers
 - Interactive Elements Determination Of Ka Lab Report Answers
- 8. Staying Engaged with Determination Of Ka Lab Report Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Determination Of Ka Lab Report Answers
- 9. Balancing eBooks and Physical Books Determination Of Ka Lab Report Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Determination Of Ka Lab Report Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Determination Of Ka Lab Report Answers
 - Setting Reading Goals Determination Of Ka Lab Report Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Determination Of Ka Lab Report Answers

- Fact-Checking eBook Content of Determination Of Ka Lab Report Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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