

Determination of K_a for a Weak Acid

Pre Lab Calculations

*Watch the provided video with example calculations before doing these problems.

Instructions: Watch the video, work these problems and upload your work and answers before your scheduled Zoom lab time. Do your best and use the video. ****MUST** show work to receive credit!

1. The pH of a 0.20 M solution of an unknown monoprotic acid, HA, is 2.12. Determine the value of the K_a for this acid. Start each problem by writing a balanced acid ionization (since calculating K_a) in water reaction and setting up an appropriate ICE table.

Step 1: Acid ionization reaction: $HA(aq) + H_2O(l) \rightleftharpoons H_3O^+(aq) + A^-(aq)$.

Step 2: Set up an ICE table

	HA	H ₂ O	A ⁻	H ₃ O ⁺
I	0.20M		0	0
C	-x		x	x
E	0.19M		$7.6 \times 10^{-3}M$	$7.6 \times 10^{-3}M$

Step 3: Determine the $[H_3O^+]_{eq}$ using the pH.

$$[H_3O^+]_{eq} = 10^{-pH} = 10^{-2.12} = 7.6 \times 10^{-3}M$$

Step 4: Fill in your known values in to the ICE table. Do $C_1V_1 = C_2V_2$ dilutions if necessary, to find concentrations. (Hint: this isn't necessary in question 1, but is necessary if you're mixing your acid with anything else).

Step 5: Write the equilibrium expression using your balanced equation from Step 1. Use equilibrium values in your ICE table to solve for K_a .

$$K_a = \frac{[H_3O^+][A^-]}{[HA]} = \frac{(7.6 \times 10^{-3}M)^2}{0.19M} = 3.04 \times 10^{-4}$$

2. A solution is prepared by adding 25.0 mL of 0.20 M HA (a different monoprotic acid than in question 1) to 50.0 mL of 0.20 M NaA. The pH of this solution is 3.45. Determine the value of K_a for this acid.

Hint: Follow the steps as outlined in problem 1. Use $C_1V_1 = C_2V_2$ to find the new initial concentration of acid after dilution with the NaA solution. The equation you used in question 1

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R Barnett



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SAM-TR. ,1972 **Table of PH, H+ and OH- Values Corresponding to Electromotive Forces Determined in Hydrogen Electrode Measurements** Carl Louis August Schmidt, Dennis Robert Hoagland, 1919 *Exercises in General, Organic, and Biological Chemistry* Arne Nels Langsjoen, 1973 **Government Reports Announcements & Index** ,1988

University of California Publications in Physiology University of California, Berkeley, 1915 Collected Reprints from the George Williams Hooper Foundation for Medical Research George Williams Hooper Foundation for Medical Research, 1918 Reprints from various medical and scientific periodicals X-Ray Fluorescence Spectroscopy for Laboratory Applications Michael Haschke, Jörg Flock, Michael Haller, 2021-04-05 Provides comprehensive coverage on using X ray fluorescence for laboratory applications This book focuses on the practical aspects of X ray fluorescence XRF spectroscopy and discusses the requirements for a successful sample analysis such as sample preparation measurement techniques and calibration as well as the quality of the analysis results X Ray Fluorescence Spectroscopy for Laboratory Applications begins with a short overview of the physical fundamentals of the generation of X rays and their interaction with the sample material followed by a presentation of the different methods of sample preparation in dependence on the quality of the source material and the objective of the measurement After a short description of the different available equipment types and their respective performance the book provides in depth information on the choice of the optimal measurement conditions and the

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