

Colorimetric Titration Experiment - Determine the amount of vitamin C in a medium peach

A suitable method for the determination of vitamin C (ascorbic acid, $C_6H_8O_6$) is a titration with potassium iodate (KIO_3). In this reaction, potassium iodate is used as a titrant and when added to an ascorbic acid solution that contains strong acid and potassium iodide (KI), the potassium iodate reacts with potassium iodide, liberating molecular iodine (I_2) as shown in the reaction below:



During the titration, as long as the solution contains ascorbic acid, the I_2 produced in equation 1 is used up in a rapid reaction with ascorbic acid (equation 2), during which dehydroascorbic acid ($C_6H_6O_6$) and iodide ion (I^-) are formed:



Once all the ascorbic acid has been consumed, any excess iodine (I_2) will remain in solution. This excess iodine reacts with starch, to form an intensely blue colored complex, indicating that the endpoint is reached.

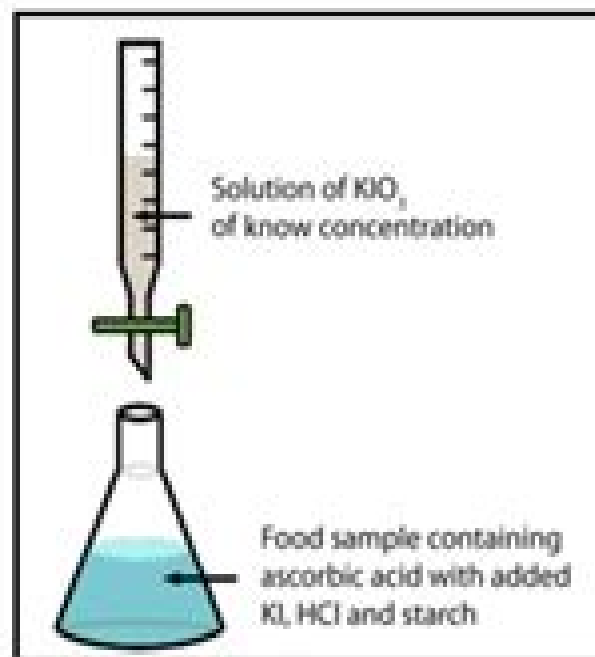
The experimental setup is shown schematically to the right.

The virtual lab stockroom contains the following solutions:

Solid potassium iodate - (KIO_3). You will need to create a stock solution with this in order to perform your titration. (Note: The molar mass of potassium iodate is 214.001 g/mol.)

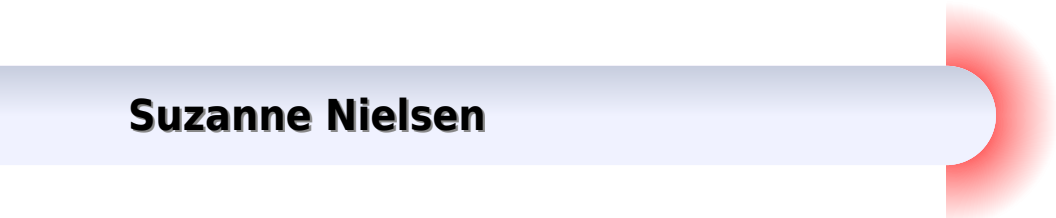
Potassium iodide solution - This is provided as a concentrated solution of KI in water. Since you are required to use an excess KI, the amount does not need to be measured precisely. Use 5 mL of KI solution for each titration.

2M HCl - The addition of acid is necessary to provide the acidic conditions required in the iodometric titration (see Equation 1). Use 4mL of the HCl solution in each titration.



Determination Of Vitamin C Concentration By Titration

Suzanne Nielsen



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Nutrition Labeling Handbook Ralph Shapiro,1995-08-30 This handbook examines the Nutritional Labeling and Education Act NLEA passed by Congress in 1990 It discusses the history of the NLEA and its impact on various segments of the food industry making complex and detailed regulations easily understandable throughout Government industry and consumer perspectives on labelling regulations are provided along with practical guidelines for compliance and packaging

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Handbook of Food Analysis: Physical characterization and nutrient analysis Leo M. L. Nollet,2004 This two volume handbook supplies food chemists with essential information on the physical and chemical properties of nutrients descriptions of analytical techniques and an assessment of their procedural reliability The new edition includes two new chapters that spotlight the characterization of water activity and the analysis of inorganic nutrients and provides authoritative rundowns of analytical techniques for the sensory evaluation of food amino acids and fatty acids neutral lipids and phospholipids and more The leading reference work on the analysis of food this edition covers new topics and techniques and reflects the very latest data and methodological advances in all chapters

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Food Analysis Suzanne Nielsen,2003-04-30 This book provides information on the techniques needed to analyze foods in laboratory experiments All topics covered include information on the basic principles procedures advantages limitations and applications This book is ideal for undergraduate courses in food analysis and is also an invaluable reference to professionals in the food industry General information is provided on regulations standards labeling sampling and data handling as background for chapters on specific methods to determine the chemical composition and characteristics of foods Large expanded sections on spectroscopy and chromatography are also included Other methods

and instrumentation such as thermal analysis selective electrodes enzymes and immunoassays are covered from the perspective of their use in the chemical analysis of foods A helpful Instructor s Manual is available to adopting professors

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